

# Chemistry 202: Quiz #9

1. Consider the reaction  $\text{NO}(g) + \text{O}_3(g) \rightarrow \text{NO}_2(g) + \text{O}_2(g)$ . You find that by doubling the partial pressures of both of the reactants, you increase the initial rate by a factor of 4. Assuming that the orders of reaction with respect to each reactant is limited to 0, 1, or 2, how many possible rate laws match this observation?
- a) 0                      b) 1                      c) 2                      d) 3                      e) 4
2. The reaction  $a\text{A} \rightarrow \text{Products}$  in which  $[\text{A}]_0 = 8.00 \text{ M}$  is 31.4% complete after 14.0 minutes and 52.9% complete after 28.0 minutes. What is the rate law for this reaction?
- a) rate =  $k$   
b) rate =  $k[\text{A}]$   
c) rate =  $k[\text{A}]^2$   
d) rate =  $k[\text{A}]^3$   
e) The rate law cannot be determined from these data.
3. We saw the decomposition of hydrogen peroxide,  $2\text{H}_2\text{O}_2(aq) \rightarrow 2\text{H}_2\text{O}(l) + \text{O}_2(g)$ , in lecture in a couple of different demonstrations. Suppose the following mechanism is proposed:
1.  $\text{H}_2\text{O}_2(aq) \rightarrow 2\text{HO}(aq)$
  2.  $\text{H}_2\text{O}_2(aq) + \text{OH}(aq) \rightarrow \text{H}_2\text{O}(l) + \text{HO}_2(aq)$
  3.  $\text{HO}_2(aq) + \text{OH}(aq) \rightarrow \text{H}_2\text{O}(l) + \text{O}_2(g)$

Which of the following best follows from this mechanism?

- a) The first step is most likely rate determining and the rate law =  $k[\text{H}_2\text{O}_2]^2$ .  
b) The first step is most likely a fast equilibrium step and the rate law =  $k[\text{H}_2\text{O}_2]^{3/2}$ .  
c) The steady state approximation is required because the rates of the steps have not been listed.  
d) The first step is most likely rate determining and the rate law =  $k[\text{H}_2\text{O}_2]$ .  
e) The first step is most likely a fast equilibrium step and the rate law =  $k[\text{H}_2\text{O}_2]$ .
4. How many of the following are true concerning the steady-state approximation?
- I. We use this approximation when the rate determining step is known.
  - II. The concentration of an intermediate is assumed to be 0  $M$ .
  - III. The rate of production of a reactant = the rate of consumption of a reactant.
  - IV. The rate law should include the intermediate given in the mechanism.
- a) 0                      b) 1                      c) 2                      d) 3                      e) 4

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5. You carry out three reactions at the same temperatures (as measured in Kelvin) and measure their rates. You then carry out these reactions at double the Kelvin temperature and find that the rate of reaction I increases by a factor of two, the rate of reaction II increases by a factor of three, and the rate of reaction III increases by a factor of four. Which reaction has the **lowest** activation energy? Assume the pre-exponential factors are constant with temperature.
- Reaction I
  - Reaction II
  - Reaction III
  - The reactions have the same activation energy.
  - We cannot determine this without knowing the temperatures.

6. Consider a reaction as represented by the general equation  $aA + bB \rightarrow \text{Products}$ .

You run this reaction with  $[A]_0 = 1.0 \times 10^{-4} M$ , and  $[B]_0 = 2.5 M$  and obtain the following data:

Time (sec)	[A] (M)
0	$1.0 \times 10^{-4}$
20.	$6.3 \times 10^{-5}$
40.	$4.0 \times 10^{-5}$
60.	$2.5 \times 10^{-5}$
80.	$1.6 \times 10^{-5}$

Assuming that the order of B limited to 0, 1, or 2, which of the following **cannot** be the value of the rate constant,  $k$  (with units of  $M$  and sec) at this temperature?

- a)  $6.8 \times 10^{-4}$     b)  $3.7 \times 10^{-3}$     c)  $9.2 \times 10^{-3}$     d)  $2.3 \times 10^{-2}$     e) All of these could be the value of  $k$ .
7. Which of the following statements concerning the kinetics demonstrations in lecture is **true**?
- We could not use the times for the reactions to determine relative initial rates for the "Iodine Clock Reaction" because the reactions happened rather quickly.
  - The catalyst in the "Activated Complex" demonstration took part in the reaction but was re-produced at the end.
  - Adding the  $MnO_2(s)$  catalyst in the "Magic Genie" demonstration produced more oxygen than otherwise could have been produced by the reaction over time.
  - Because the concentrations of both HCl and acetic acid were  $3.0 M$ , they reacted with baking soda with the same rates.
  - None of these are true.

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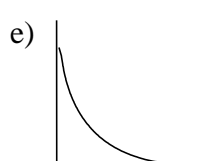
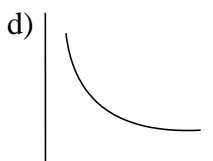
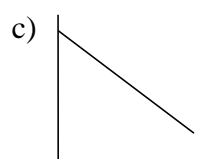
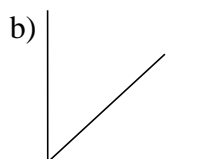
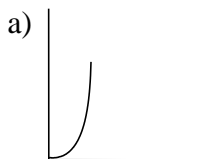
8. Consider the decomposition of ethanol,  $\text{C}_2\text{H}_5\text{OH}(g) \rightarrow \text{C}_2\text{H}_4(g) + \text{H}_2\text{O}(g)$ , at  $125^\circ\text{C}$ . You run the reaction and collect the following data:

Time (sec)	P[C <sub>2</sub> H <sub>5</sub> OH] (atm)
0	3.200
30.00	1.684
60.00	1.143
90.00	0.8649
120.0	0.6957

What is the order of the reaction with respect to ethanol?

- a) Zero order
  - b) First order
  - c) Second order
  - d) It cannot be determined with these data.
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- 9, 10. Choose the best graph for the plots described below.

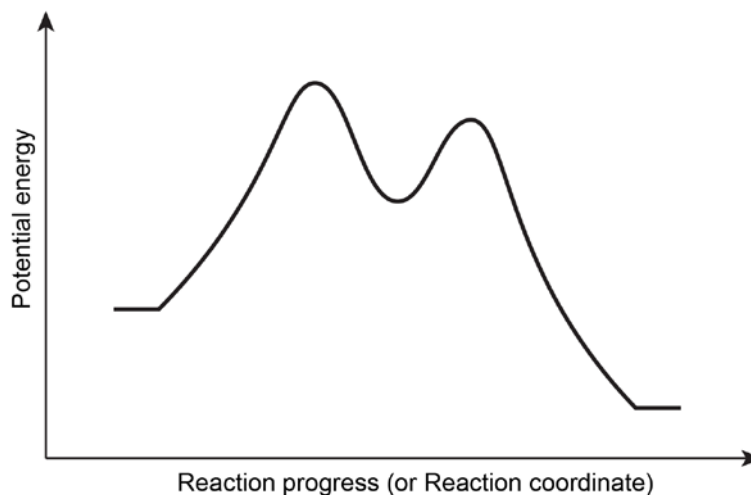


9.  $\ln(k)$  vs.  $1/T$  (K) for a typical chemical reaction.
10. A plot of  $[A]$  vs. time for reaction type  $aA \rightarrow \text{Products}$  which is not zero-order in A.

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11. Let's consider the reaction in which ozone becomes oxygen gas [ $2\text{O}_3(g) \rightarrow 3\text{O}_2(g)$ ]. Consider you are doing a kinetics experiment with this reaction and you have the following information:

- In monitoring the reaction, you find the successive half-lives to be constant.
- The activation energy of the reaction is about 300 kJ.
- The reaction coordinate looks like the following:



In this problem you will be developing a mechanism, evaluating experimental data, and applying the steady-state approximation.

Please **show all work** and **provide concise, yet complete, explanatory answers**. Think about what you want to do/say before writing.

- a. **Develop a mechanism** for the reaction and **determine the rate law**. **Defend your answer** by explaining how you used each of the three pieces of information given on the previous page. In your discussion, use the bond energies on Table 13.6 (you can also use your estimates of  $\Delta H^\circ_f$  for ozone and atomic oxygen from problem #21 if you wish) to **justify the shape of the reaction coordinate**. Also, **discuss** how your mechanism and rate law **make sense chemically**.

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- b. You are going to run the experiment starting with 5.000 atm  $O_3$  in an otherwise evacuated 10.0-L rigid container at 298K. The first two readings are given. Please **fill in the rest of the expected data** in the table below and **defend your answers** with calculations. Note: you are to estimate as closely as possible the time when maximum pressure is reached and include it in the table below. **Show/explain all work.**

Time (seconds)	$P_{\text{total}}$ (atm)
0	5.000
15.00	5.732
30.00	
45.00	
60.00	
75.00	
	Max. =

- c. You should know which step is the rate-determining-step both from the reaction coordinate and from your mechanism (and you should have discussed this in part a). But just to make sure, you decide to **apply the steady-state approximation (SSA)** to your mechanism. **Define the rate** of the reaction in **two ways**, and show how the SSA in both cases **gives consistent answers** to each other. (use different symbols for the rate constants as we did in lecture, text, and videos, such as  $k_1$ , etc. to distinguish the different steps of the mechanism as appropriate).

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**KEY:**

MC: 1. d, 2. b, 3. d, 4. a, 5. a, 6. a, 7. b, 8. c, 9. c, 10. e

11. See videos, lectures, and the textbook.