

Chapter Six + Seven Homework

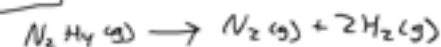
Chapter 6-6

- color change typically indicates a chemical change

6-13



6-16

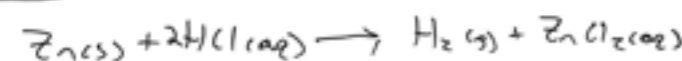


6-18



remember!
 $\begin{cases} \text{Ag}^+ \text{O}^{2-} \\ \text{Ag}_2^+ \end{cases}$

6-24

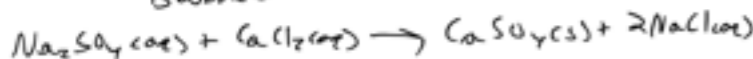


6-40

a) Unbalanced



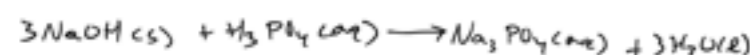
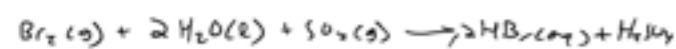
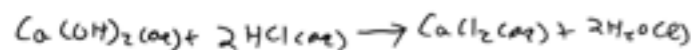
Balanced



b) Unbalanced



Balanced



6-48 **Rxn 1**



Rxn 2



or



6-52

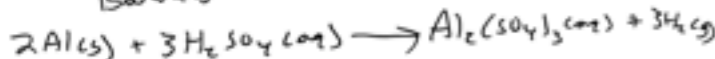
Fractions can be used to help solve the problem, BUT the best solution has only whole numbers

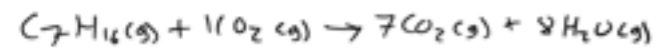
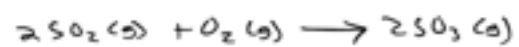
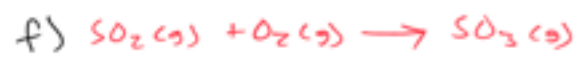
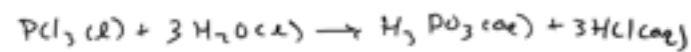
6-73

d) Unbalanced



Balanced

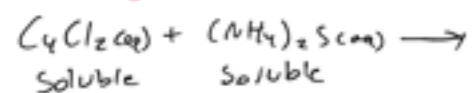




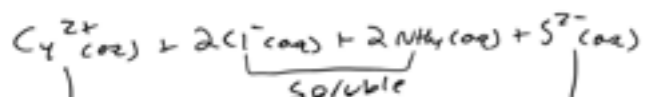
7-11

PbS	Insoluble	Rule #6
Mg(OH) ₂	Insoluble	#5
Na ₂ SO ₄	Soluble	#2
(NH ₄) ₂ S	Soluble	#2
BaCO ₃	Insoluble	#6
AlPO ₄	Insoluble	#6
PbCl ₂	Insoluble	#3 Exception
CaSO ₄	Insoluble	#4 Exception

7-15 (A)



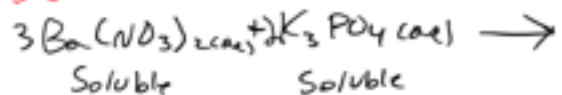
Soluble Soluble



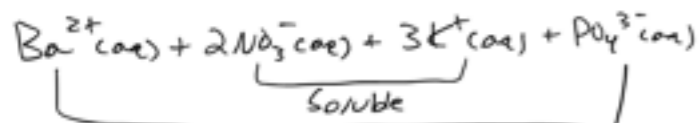
Insoluble



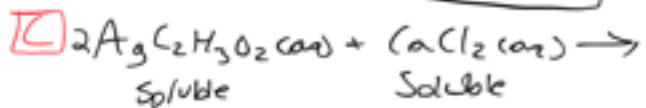
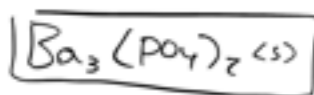
(B)



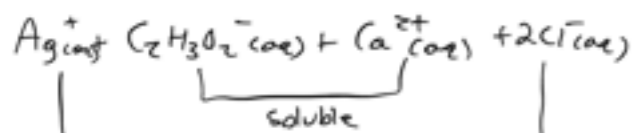
Soluble Soluble



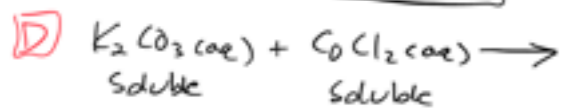
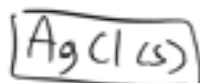
Insoluble



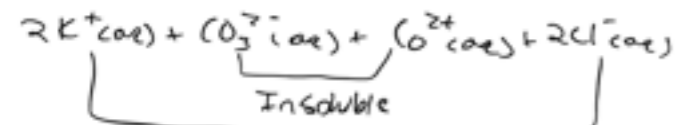
Soluble Soluble



Insoluble



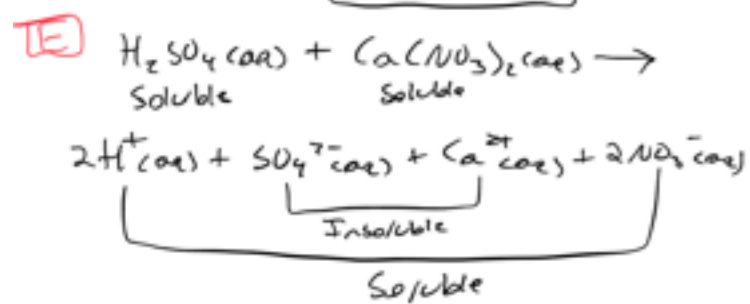
Soluble Soluble



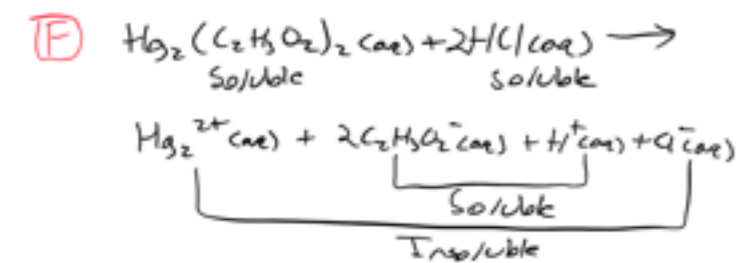
Insoluble

Soluble

Co (CO₃ cs)

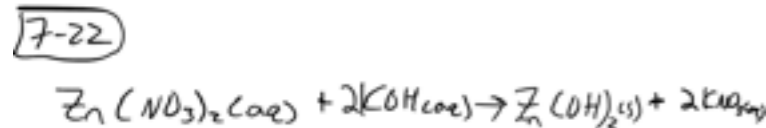
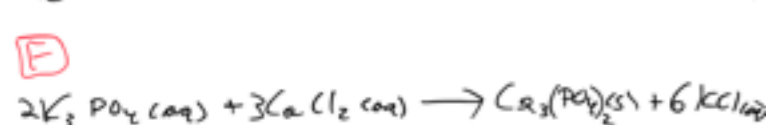
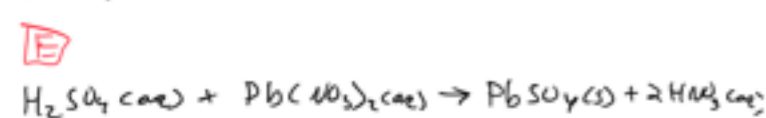
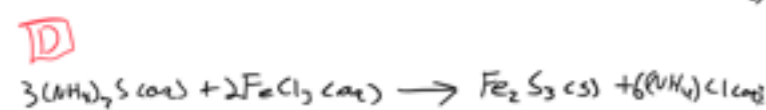
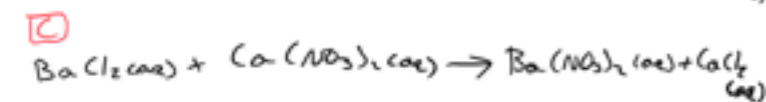
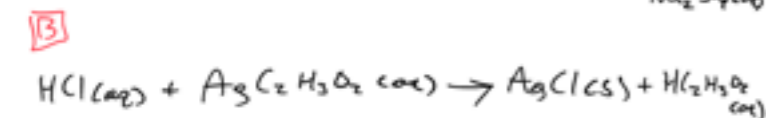
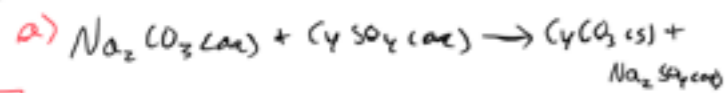


Ca SO₄ cs

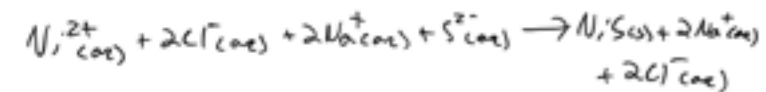
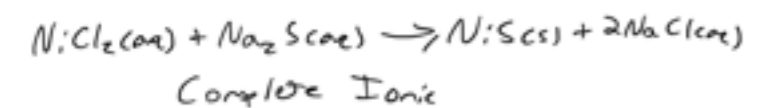


Hg₂ Cl₂ cs

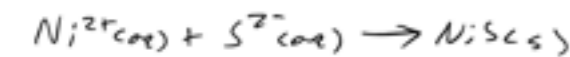
7-18 Write Balanced Molecular Formula



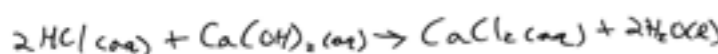
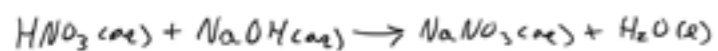
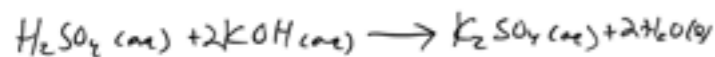
7-26 Balanced Molecular

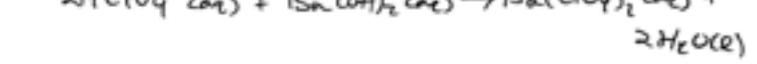


Net Ionic



7-40





7-75 Balanced Net Ionic

